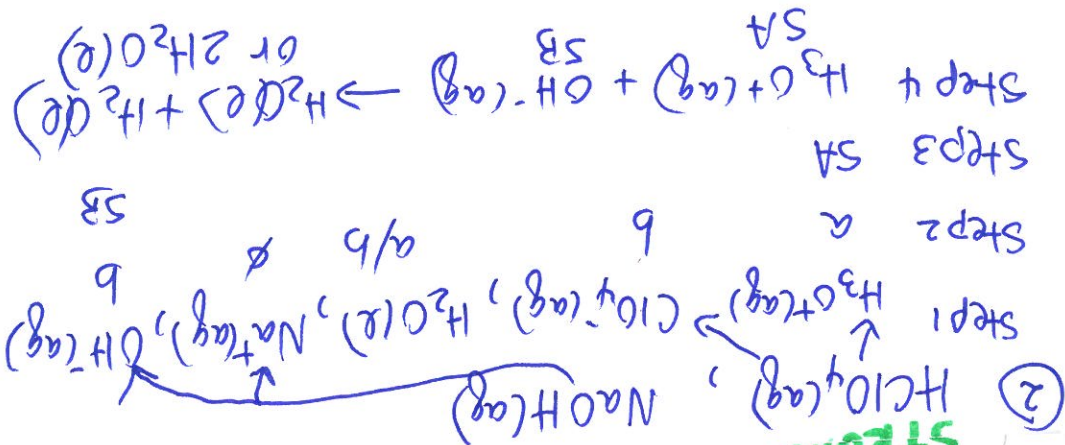


Use the five step approach and an acid/base table to predict the acid base reaction for the following:

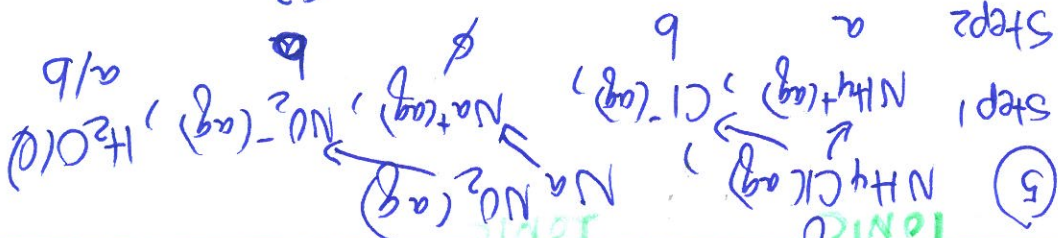
1. Methanoic acid is added to an aqueous solution of sodium hydrogen sulfate.
2. Strong acids, such as perchloric acid, have been shown to react quantitatively with strong bases, such as sodium hydroxide.
3. Predict the acid-base reaction of bleach, sodium hypochlorite, with vinegar.
4. Hydrofluoric acid and an aqueous solution of sodium hydrogen sulfate are mixed to test the five step method of predicting acid-base reactions.
5. A student mixes solutions of ammonium chloride and sodium nitrite in a chemistry laboratory.
6. Empirical lab analysis has shown that nitric acid reacts quantitatively with a sodium acetate solution.
7. A consumer attempts to neutralize an aqueous sodium hydrogen sulfate cleaner with a solution of lye. (lye is also known as sodium hydroxide)
8. Can ammonium nitrate fertilizer, added to water, be used to neutralize a muriatic acid spill? (muriatic acid is also known as hydrochloric acid)

STRONG



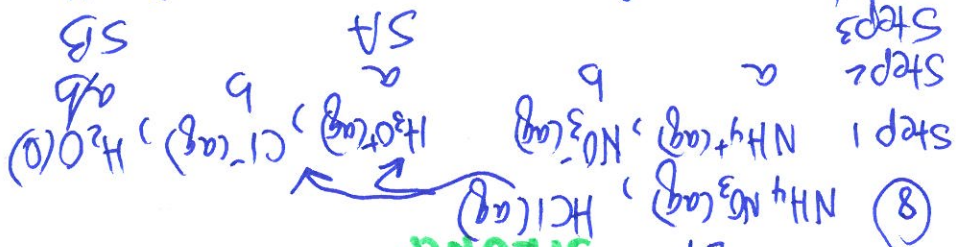
Steps you were given so →

IONIC



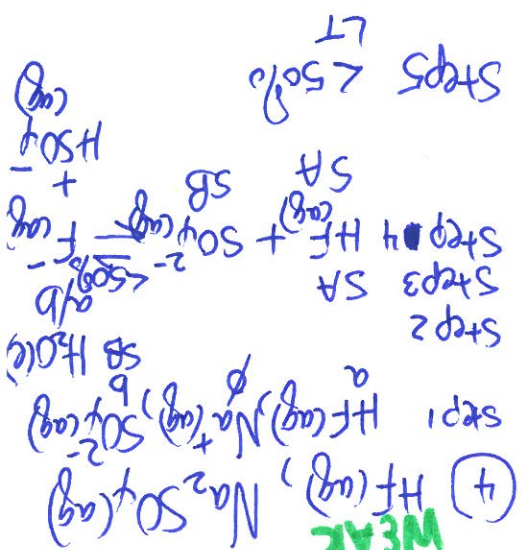
Steps < 50% since SA is below SB in acid base table

STRONG



Steps same products!! Not a good choice

WEAK



WEAK

